

Helpful Hints:

How to recognize a polyatomic ion:

- 2 or more elements covalently bonded carrying a charge

- majority of names end in "ate" or "ite"

- "ate" always has 1 more oxygen than "ite"

- "ate" + "ite" generally means oxygen is involved

- per = ate "per" = 1 more oxygen than "ate"

- hypo = ite "hypo" = 1 less oxygen than "ite"

SO_5^{2-} - persulfate

ates tend to be $-O_3$
ites tend to be $-O_4$

					VIIIA
					2
13	14	15	16	17	He
IIIB	IVB	VB	VIB	VII B	helium
IIIA	IVA	VA	VIA	VIIA	4.003
5 B boron 10.81	6 C carbon 12.01	7 N nitrogen 14.01	8 O oxygen 16.00	9 F fluorine 19.00	10 Ne neon 20.18
13 Al aluminum 26.98	14 Si silicon 28.09	15 P phosphorus 30.97	16 S sulfur 32.07	17 Cl chlorine 35.45	18 Ar argon 39.95
31 Ga gallium 69.72	32 Ge germanium 72.61	33 As arsenic 74.92	34 Se selenium 78.96	35 Br bromine 79.90	36 Kr krypton 83.80
49 In indium 114.8	50 Sn tin 118.7	51 Sb antimony 121.8	52 Te tellurium 127.6	53 I iodine 126.9	54 Xe xenon 131.3
81 Tl thallium 204.4	82 Pb lead 207.2	83 Bi bismuth 209.0	84 Po polonium (209)	85 At astatine (210)	86 Rn radon (222)