

## Calculating Molar Mass #2

Calculate the Molar Mass of the following.  
Show your work.

1.) Hypochlorous acid  $\text{HClO}$   
 $1.01 + 35.45 + 16$   
 $= 52.46 \text{ g/mol}$

2.) Copper (I) phosphate  $\text{Cu}_3\text{PO}_4$   
 $(63.55 \times 3) + 30.97 + (16 \times 4)$   
 $= 285.62 \text{ g/mol}$

3.) Hydrocyanic acid  $\text{HCN}$   
 $1.01 + 12.01 + 14.01$   
 $= 27.03 \text{ g/mol}$

4.) Potassium oxide  $\text{K}_2\text{O}$   
 $(2 \times 39.1) + 16$   
 $= 94.2 \text{ g/mol}$

5.) Beryllium oxide  $\text{BeO}$   
 $9.012 + 16$   
 $= 25.01 \text{ g/mol}$

6.) Pentanitrogen decoxide  $\text{N}_5\text{O}_{10}$   
 $(5 \times 14.01) + (10 \times 16)$   
 $= 230.05 \text{ g/mol}$

7.) Chromic acid  $\text{H}_2\text{CrO}_4$   
 $(2 \times 1.01) + 52 + (4 \times 16)$   
 $= 118.02 \text{ g/mol}$

8.) Copper (II) sulfide  $\text{CuS}$   
 $63.55 + 32.07 = 95.62 \text{ g/mol}$

9.) Sulfur trioxide  $\text{SO}_3$   
 $32.07 + (3 \times 16) = 80.07 \text{ g/mol}$

10.) Zinc carbonate  $\text{ZnCO}_3$   
 $65.39 + 12.01 + (3 \times 16) = 125.38 \text{ g/mol}$

11.) Iodous acid  $\text{HIO}_2$   
 $1.01 + 126.9 + (2 \times 16)$   
 $= 159.91 \text{ g/mol}$

12.) Nitrogen dioxide  $\text{NO}_2$   
 $14.01 + 2 \times 16 = 46.01 \text{ g/mol}$

13.) Iron (II) citrate  $\text{Fe}_3(\text{C}_6\text{H}_5\text{O}_7)_2$   
 $(3 \times 55.85) + (12 \times 12.01) + (10 \times 1.01)$   
 $+ (14 \times 16) = 545.77 \text{ g/mol}$

14.) Hydrothiocyanic acid  $\text{HSCN}$   
 $1.01 + 32.07 + 12.01 + 14.01$   
 $= 59.01 \text{ g/mol}$

15.) Diphosphorous trioxide  $\text{P}_2\text{O}_3$   
 $2 \times 30.97 + 3 \times 16$   
 $= 109.94 \text{ g/mol}$