

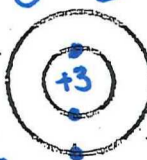
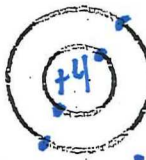
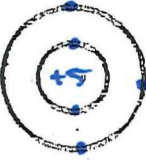


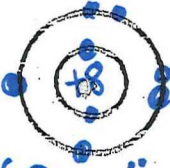

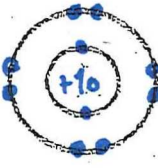
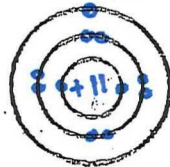
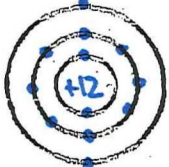
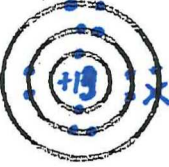
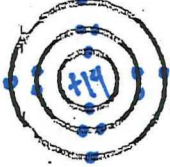
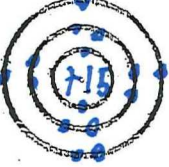
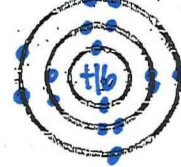
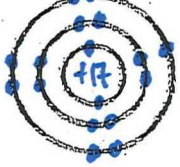
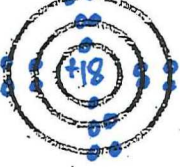


ATOMIC NUMBER/MASS NUMBER/BOHR MODEL/LEWIS DOT D. GRAM REVIEW

1. WRITE THE NAME OF EACH ELEMENT IN THE BOX.
2. Using the short PTE write the group numbers in Roman Numerals
3. Label the periods.
4. If a group has a specific name, write it about the group
5. Put the correct nuclear charge inside the nucleus.
6. In each box include the symbol, atomic number and mass number.
7. Draw the Bohr Model for each element.
8. Draw the Lewis Dot Diagram for each element.

	I alkali metals	Alkaline Earth Metals II	III	IV	V	VI Chalcogens	Halogens VII	Noble Gases VIII
1	 ${}^1_1\text{H}$ H hydrogen							 ${}^2_2\text{He}$ He Helium
2	 ${}^7_3\text{Li}$ Li Lithium	 ${}^9_4\text{Be}$ Be Beryllium	 ${}^5_5\text{B}$ B Boron	 ${}^{12}_6\text{C}$ C Carbon	 ${}^{14}_7\text{N}$ N Nitrogen	 ${}^{16}_8\text{O}$ O Oxygen	 ${}^{19}_9\text{F}$ F Fluorine	 ${}^{20}_{10}\text{Ne}$ Ne Neon
3	 ${}^{23}_{11}\text{Na}$ Na Sodium	 ${}^{24}_{12}\text{Mg}$ Mg Magnesium	 ${}^{27}_{13}\text{Al}$ Al Aluminum	 ${}^{28}_{14}\text{Si}$ Si Silicon	 ${}^{31}_{15}\text{P}$ P Phosphorus	 ${}^{32}_{16}\text{S}$ S Sulfur	 ${}^{35}_{17}\text{Cl}$ Cl Chlorine	 ${}^{40}_{18}\text{Ar}$ Ar Argon