

Wed 1/8

NOMENCLATURE PRACTICE:

EX: calcium hydroxide Ca(OH)_2

1. sodium fluoride $\begin{matrix} +1 & -1 \\ \text{NaF} \end{matrix}$
2. ammonium sulfide $\begin{matrix} +1 & -2 \\ (\text{NH}_4)_2\text{S} \end{matrix}$
3. magnesium phosphide $\begin{matrix} +2 & -3 \\ \text{Mg}_3\text{P}_2 \end{matrix}$
4. aluminum hydroxide $\begin{matrix} +3 & -1 \\ \text{Al(OH)}_3 \end{matrix}$
5. calcium chloride $\begin{matrix} +2 & -1 \\ \text{CaCl}_2 \end{matrix}$
6. ammonium oxide $\begin{matrix} +1 & -2 \\ (\text{NH}_4)_2\text{O} \end{matrix}$
7. potassium carbonate $\begin{matrix} +1 & -2 \\ \text{K}_2\text{CO}_3 \end{matrix}$
8. barium phosphate $\begin{matrix} +2 & -3 \\ \text{Ba}_3(\text{PO}_4)_2 \end{matrix}$
9. sodium nitride $\begin{matrix} +1 & -3 \\ \text{Na}_3\text{N} \end{matrix}$
10. magnesium sulfate $\begin{matrix} +2 & -2 \\ \text{MgSO}_4 \end{matrix}$
11. barium sulfide $\begin{matrix} +2 & -2 \\ \text{BaS} \end{matrix}$
12. ammonium phosphide $\begin{matrix} +1 & -3 \\ (\text{NH}_4)_3\text{P} \end{matrix}$
13. potassium nitrate $\begin{matrix} +1 & -1 \\ \text{KNO}_3 \end{matrix}$
14. aluminum oxide $\begin{matrix} +3 & -2 \\ \text{Al}_2\text{O}_3 \end{matrix}$
15. sodium carbonate $\begin{matrix} +1 & -2 \\ \text{Na}_2\text{CO}_3 \end{matrix}$
16. magnesium hydroxide $\begin{matrix} +2 & -1 \\ \text{Mg(OH)}_2 \end{matrix}$
17. potassium sulfide $\begin{matrix} +1 & -2 \\ \text{K}^+ \text{S}^{2-} \rightarrow \text{K}_2\text{S} \end{matrix}$
18. barium chloride $\begin{matrix} +2 & -1 \\ \text{BaCl}_2 \end{matrix}$
19. aluminum phosphate $\begin{matrix} +3 & -3 \\ \text{AlPO}_4 \end{matrix}$
20. calcium fluoride $\begin{matrix} +2 & -1 \\ \text{CaF}_2 \end{matrix}$

CRITICAL THINKING:

1. When a compound is balanced, what is the relationship between the positive and negative oxidation numbers?
2. What is the significance of a compound whose name ends in "ide"?
3. What is the oxidation number trend for the elements in the short periodic table?